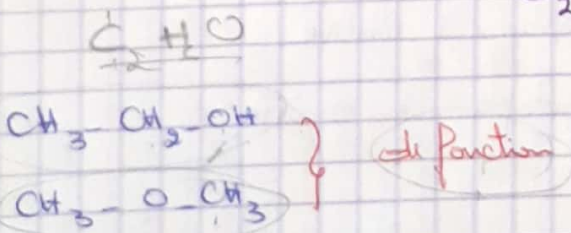
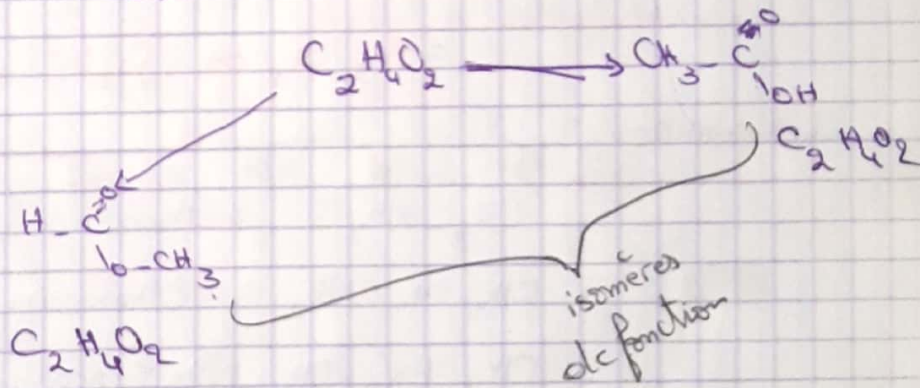
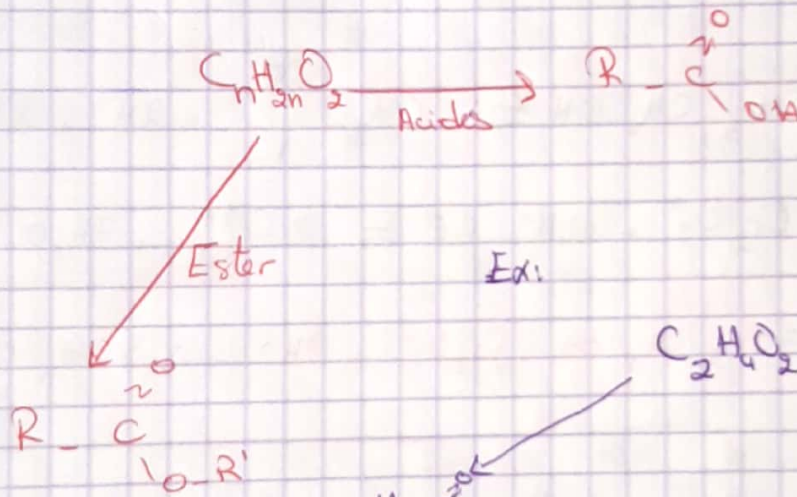
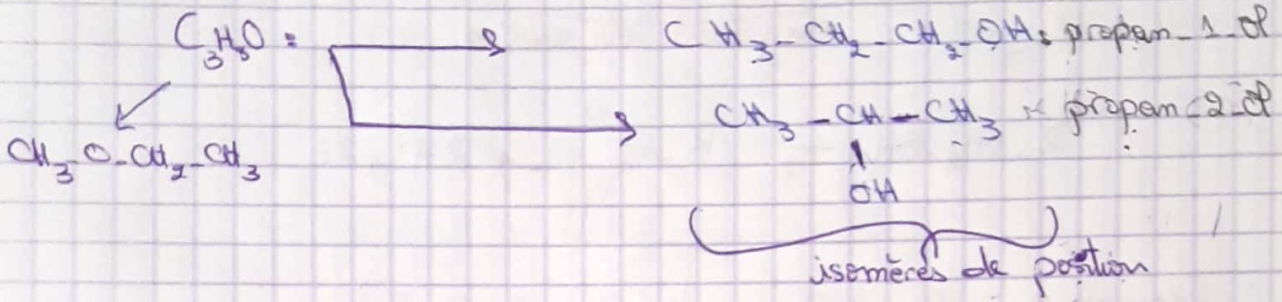
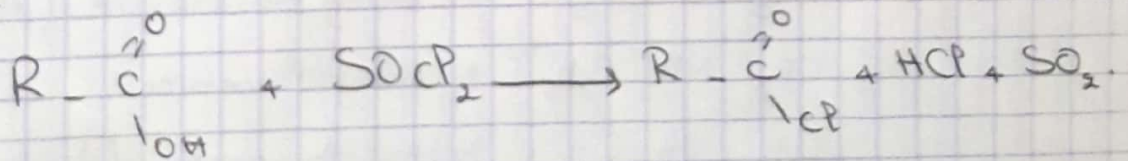


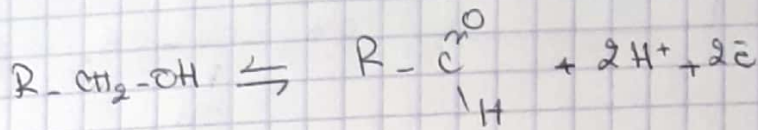
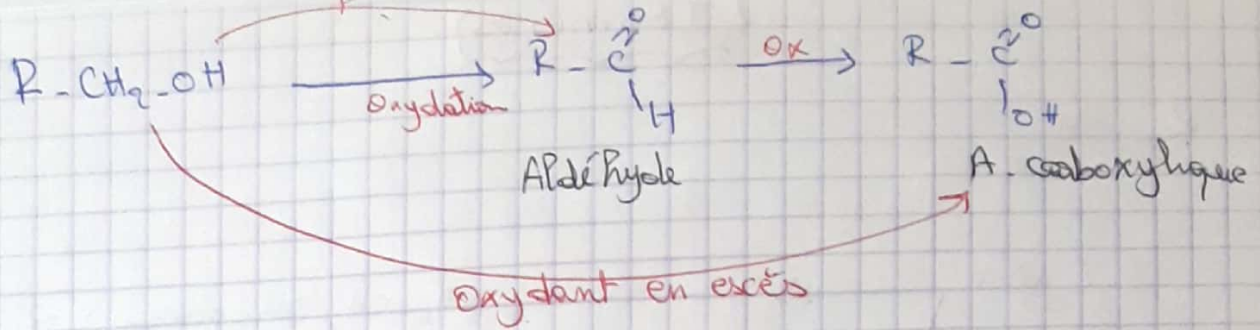
} Isomères }

R-OH

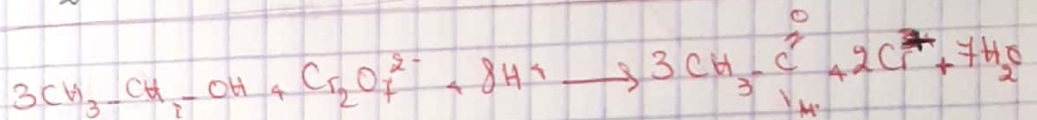
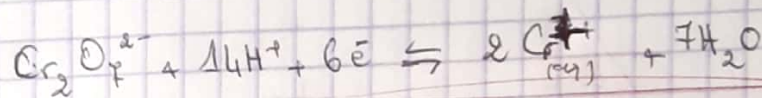
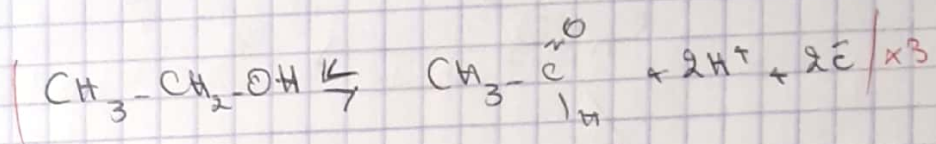


C_2H_4O

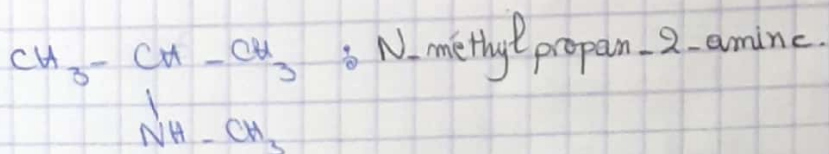
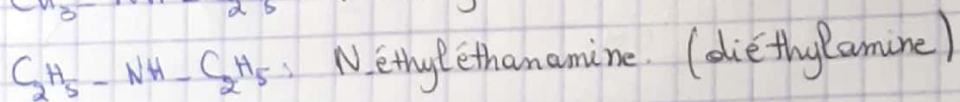
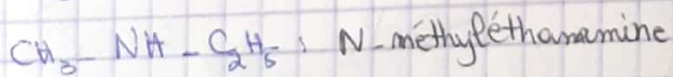
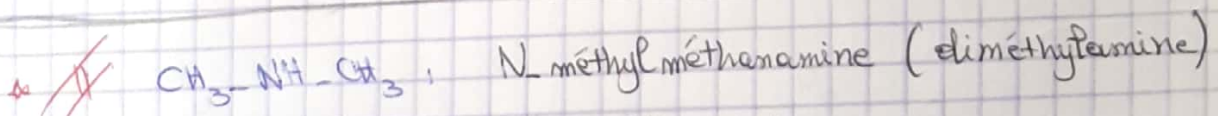
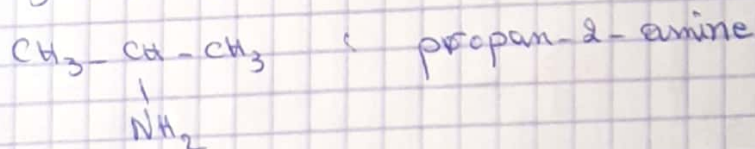
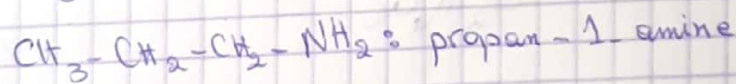
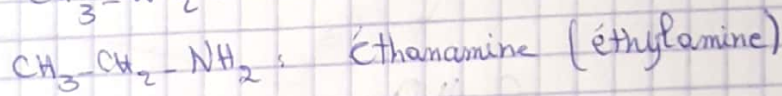
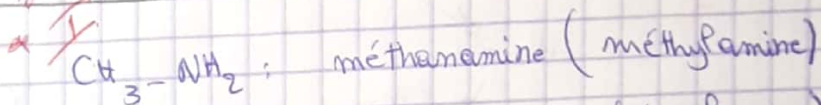




Exemple : Oxydation de l'éthanol par $Cr_2O_7^{2-}$



Les amines :



Prisme

* Cas particulier, $\textcircled{1}$

"petits angles, $\sin d \approx d$

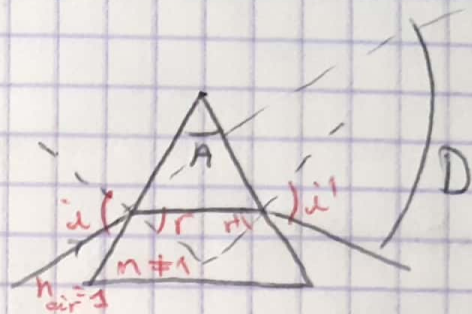
$$i = nr$$

$$i' = nr'$$

$$D = nr + nr' - A$$

$$= n(r + r') - A$$

$$D \approx A(n - 1) \approx 19,5^\circ$$



$$\sin i = n \sin r$$

$$\sin i' = n \sin r'$$

$$A = r + r'$$

$$D = i + i' - A$$

Cas particulier ② :

$$i = i' \Rightarrow D = D_{\min}$$



$$r = r' \Rightarrow A = 2r = 2r'$$



$$r = r' = \frac{A}{2}$$

$$D_{\min} = i + i' - A = 2i - A$$



$$i = \frac{D_{\min} + A}{2}$$

$$n = \frac{\sin\left(\frac{D_{\min} + A}{2}\right)}{\sin\left(\frac{A}{2}\right)}$$

cas particulier ③

"incidence normal"

$$i = 0$$

$$r' = 0$$

$$A = r'$$

$$\sin i' = n \sin r' \Rightarrow i' = \sin^{-1}(n \sin A)$$

$$D = \sin^{-1}(n \sin A) - A$$

cas particulier ④:

"Emergence normal"

$$i' = 0$$

$$r = 0$$

$$A = r$$

$$A = r$$

$$A = r$$

$$\sin i = n \sin A$$

$$D = \sin^{-1}(n \sin A) - A$$

base forte \Rightarrow

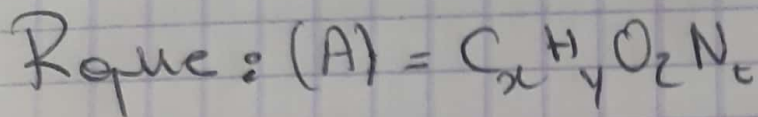
$$pH = 14 + \log C_B$$

Base Faible \Rightarrow $pH < 14 + \log C_B$

Acide fort \Rightarrow

$$pH = -\log C_A$$

Acide Faible \Rightarrow $pH > -\log C_A$



$$\%C = \frac{xM(C)}{M(A)} \times 100$$

$$\%H = \frac{yM(H)}{M(A)}$$

$$\%O = \frac{zM(O)}{M(A)} \times 100$$

$$\frac{M(A)}{100} = \frac{xM(C)}{\%C} = \frac{yM(H)}{\%H}$$

$$= \frac{zM(O)}{100}$$